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Stoichiometry the calculation of quantities in chemical reactions  
Mole ratio a conversion factor derived from the coefficients of a  
balanced chemical equation interpreted in terms of moles  
Limiting reagent

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1 Chapter 12: Stoichiometry Test Review 1. Definitions of:

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theoretical yield, percent yield, stoichiometry, mole ratio, limiting reactant, excess reactant Chapter 12 Stoichiometry Practice Problems Answer ...

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a. add the coefficients of the reagents  
c. convert given quantities to volumes  
b. convert given quantities to moles  
d ...

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Download Free Chapter 12 Stoichiometry Answer Key Pearson Chapter 12 Stoichiometry Test Answer Key Answer:  $4.93 \times 10^{-5}$  L or 49.3  $\mu$ L In Example 12.2.1 and Example 12.2.2, the identity of the limiting reactant has been apparent:  $[\text{Au}(\text{CN})_2]^-$ ,  $\text{LaCl}_3$ , ethanol, and para -nitrophenol. When

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- a. add the coefficients of the reagents
- c. convert given quantities to volumes
- b. convert given quantities to moles
- d. convert given

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quantities to masses \_\_\_\_ 2.

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